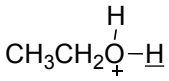
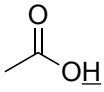
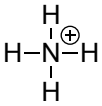
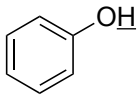
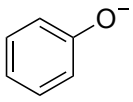
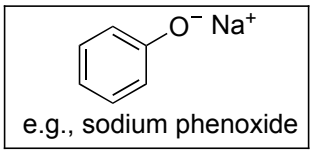
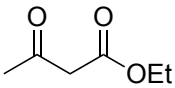
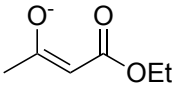
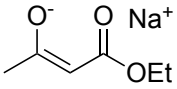
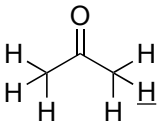
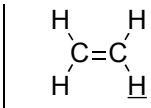
pKa Table: Effect of electronegativity and resonance

Acidity

Acid	pKa	Conjugate Base	
HCl	-7	Cl^-	
H_2SO_4	-4	HSO_4^-	
HNO_3	-2	NO_3^-	
	-2	CH_3CH_2OH	
	-2	H_2O	
	5		
	9 (10)	H_3N	
 phenol	10	 phenoxide	 e.g., sodium phenoxide
 acetoacetic ester	11	 [only most stable resonance shown here]	e.g. 
HO-H water	15.7 (15)	HO^- hydroxide	$HO^-\ Na^+$ e.g., sodium hydroxide
CH_3CH_2OH ethanol	16 (15)	$CH_3CH_2O^-$ ethanoxide	$CH_3CH_2O^-\ Na^+$ e.g., sodium ethoxide
 acetone	20	 enolate [only most stable resonance shown here]	 e.g., lithium enolate
H−C≡C−H alkyne	25	$H-C\equiv C^{\ominus}$ acetylide	$H-C\equiv C\ Na^+$ e.g., sodium acetylide
H_2N-H ammonia	35	H_2N^- amide	$H_2N^-\ Na^+$ e.g., sodium amide
H-H Hygrogen	35	H^- Hydride	$H^-\ Na^+$ e.g., sodium hydride



alkene

44

$CH_3CH_2CH_2CH_2$-H

Butane

60

$CH_3CH_2CH_2CH_2^-$

Butyl carbanion

$CH_3CH_2CH_2CH_2^-$ Li^+

e.g., butyl lithium

Basicity